

Covalent

Prop.

- 1) shares electrons
- 2) 2 nonmetals

Naming

→ prefix name, ide
 PREFIX ROOT, ide
 ↓

xy

EX. CCl₄

Carbon tetrachloride

Polarity:



Lewis Structures

- inventory
- skeleton
- share w/ neighbor
- C₂H₃O₂ → acetate
- extra e⁻ on center

Ionic — transfer of electrons

Prop.

- 1) bulk crystal
- 2) dissolve in water
- 3) cation + anion
m nm
- 4) high melting point
- 5) dissolving
- breaks apart in solution

EX. NaCl → Na⁺ + Cl⁻

Melting Points:

- difference in charge conducts
 $F = \frac{q^+ q^-}{d^2} \rightarrow$ larger electricity

- Naming ↑ melting point

- Name of cation, Name of anion

- NEVER gives quantities

- ide (negative ending)

EX: NaCl Sodium chloride

* picture of dissolving water